

Stoichiometry Lab Experiment

You are working for a company that makes water-softening agents for homes with hard water. Recently, there was a mix-up on the factory floor, and sodium carbonate solution was mistakenly mixed in a 575 L container with an unknown quantity of distilled water. You must determine the amount of Na_2CO_3 in the container in order to properly predict the percentage yield of the water-softening product.

This procedure involves a double-displacement reaction between strontium nitrate, $\text{Sr}(\text{NO}_3)_2$, and sodium carbonate, Na_2CO_3 . In general, this reaction can be used to determine the amount of any carbonate compound in a solution. The solid formed will be the strontium product.

Remember that accurate results depend on precise mass measurements. Keep all glassware very clean, and do not lose any reactants or products during your lab work.

You will react an unknown amount of sodium carbonate with an excess of strontium nitrate. After purifying the product, you will determine the following:

- how much product is present
- how much Na_2CO_3 must have been present to produce that amount of product
- how much Na_2CO_3 is contained in the 575 L of solution

Procedure

Day 1

1. Measure the mass of a piece of filter paper to the nearest 0.01 g and record.
2. Set up a filtering apparatus. (funnel, 125 mL flask and filter paper)
3. Determine the mass of a clean, dry **100 mL** beaker.
4. Measure 15 mL of the Na_2CO_3 solution into the graduated cylinder. Pour the Na_2CO_3 solution into a clean, empty **50 mL** beaker. Carefully wash the graduated cylinder, and rinse it with distilled water.
5. Measure 25 mL of the 0.30 M $\text{Sr}(\text{NO}_3)_2$ solution into the graduated cylinder. Pour the $\text{Sr}(\text{NO}_3)_2$ solution into the **50 mL** beaker with the Na_2CO_3 solution. Gently stir the solution and precipitate with a glass stirring rod.
6. Slowly pour the mixture into the funnel. Be careful not to overfill the funnel because some of the precipitate could be lost between the filter paper and the funnel.
7. Rinse the beaker several more times with small amounts of distilled water. Pour the rinse water into the funnel each time.
8. After all of the solution and rinses have drained through the funnel; slowly rinse the precipitate on the filter paper in the funnel with distilled water to remove any soluble impurities.
9. Carefully remove the filter paper from the funnel, and place it with the precipitate in the **100 mL** beaker from step 3. Then place the beaker in the drying oven overnight.

Day 2

10. Measure and record the mass of the beaker with the filter paper, and precipitate to the nearest 0.01 g.
11. Dispose of the precipitate in a designated waste container.

Experiment 6 Stoichiometry Lab Report Conclusion

Jessica J Manson



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