

# Ideal Gas Law

## Chem Worksheet 14-4

Name \_\_\_\_\_

The ideal gas law is an equation that relates the volume, temperature, pressure and amount of gas particles to a constant. The ideal gas constant is abbreviated with the variable  $R$  and has the value of  $0.0821 \text{ atm}\cdot\text{L}/\text{mol}\cdot\text{K}$ . The ideal gas law can be used when three of the four gas variables are known. When using this equation it is important that the units for **pressure** are atmospheres (atm), **volume** is in liters (L), and **temperature** is converted to kelvins (K). The **amount** of gas is measured in units called moles (mol).

### USEFUL EQUATIONS

$$PV = nRT$$

$$R = 0.0821 \frac{\text{atm}\cdot\text{L}}{\text{mol}\cdot\text{K}}$$

$$T_K = T_C + 273$$

$$1 \text{ cm}^3 = 1 \text{ mL}$$

$$1 \text{ L} = 1000 \text{ mL}$$

$$1.00 \text{ atm} = 101300 \text{ Pa}$$

$$1.00 \text{ atm} = 101.3 \text{ kPa}$$

$$1.00 \text{ atm} = 760 \text{ mmHg}$$

$$1.00 \text{ atm} = 760 \text{ torr}$$

$$1.00 \text{ atm} = 14.7 \text{ psi}$$

Unknown	Equation	Known Variables
pressure	$P = \frac{nRT}{V}$	amount, temp., volume
volume	$V = \frac{nRT}{P}$	amount, temp., pressure
temperature	$T = \frac{PV}{nR}$	pressure, volume, amount
amount	$n = \frac{PV}{RT}$	pressure, volume, temp.

### example

The pressure exerted by 2.8 moles of argon gas at a temperature of  $85^\circ\text{C}$  is 420 torr. What is the volume of this sample?

- list the variables:  $P = 420 \text{ torr}$      $V = ?$      $n = 2.8 \text{ mol}$      $R = 0.0821 \frac{\text{atm}\cdot\text{L}}{\text{mol}\cdot\text{K}}$      $T = 85^\circ\text{C}$
- convert the variables:  $\frac{420 \text{ torr}}{1} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 0.553 \text{ atm}$      $T = 85^\circ\text{C} + 273 = 358 \text{ K}$
- substitute into the equation:  $V = \frac{nRT}{P}$      $V = \frac{(2.8 \text{ mol})(0.08206 \frac{\text{atm}\cdot\text{L}}{\text{mol}\cdot\text{K}})(358 \text{ K})}{0.553 \text{ atm}} = 82 \text{ L}$

### Solve the following problems.

1. A tank contains 115 moles of neon gas. It has a pressure of 57 atm at a temperature of  $45^\circ\text{C}$ . Calculate the volume of the tank.
2. A scuba tank has a pressure of 195 atm at a temperature of  $10^\circ\text{C}$ . The volume of the tank is 350 L. How many moles of air are in the tank?
3. A helium-filled balloon has a volume of 208 L and it contains 9.95 moles of gas. If the pressure of the balloon is 1.26 atm, determine the temperature in Celsius degrees.
4. A tank of oxygen has a volume of 1650 L. The temperature of the gas inside is  $35^\circ\text{C}$ . If there are 9750 moles of oxygen in the tank what is the pressure in PSI?
5. A canister of acetylene has a volume of 42 L. The temperature of the acetylene is 305 K and the pressure is 780 torr. Determine the amount (moles) of gas in the canister.
6. Calculate the volume of a  $\text{CO}_2$  cartridge that has a pressure of 850 PSI at a temperature of  $21^\circ\text{C}$ . The cartridge contains 0.273 mol of  $\text{CO}_2$ .
7. A tank contains 2500 L of argon gas. The pressure is 13790 kPa and the temperature is  $25^\circ\text{C}$ . How many moles of argon are in the tank?

# Ideal Gas Law Chem 14 4 Answers

**Anthony S. Fauci**



## **Ideal Gas Law Chem 14 4 Answers:**

**Chemical Thermodynamics at a Glance** H. Donald Brooke Jenkins, 2008-04-30 Chemical thermodynamics considers the energy transformations which drive or which occur as a result of chemical reactions It is a central discipline of chemistry and chemical engineering allowing prediction of the direction of spontaneous chemical change and the position of chemical equilibrium in any reacting system Being grounded in maths it is often perceived as a difficult subject and many students are never fully comfortable with it Chemical Thermodynamics at a Glance provides a concise overview of the main principles of Chemical Thermodynamics for students studying chemistry and related courses at undergraduate level Based on the highly successful and student friendly at a Glance approach the information is presented in integrated self contained double page spreads of text and illustrative material The material developed in this book has been chosen to ensure the student grasps the essence of thermodynamics so those wanting an accessible overview will find this book an ideal source of the information they require In addition the structured presentation will provide an invaluable aid to revision for students preparing for examinations

Physical Chemistry Horia Metiu, 2006-02-21 This is a new undergraduate textbook on physical chemistry by Horia Metiu published as four separate paperback volumes These four volumes on physical chemistry combine a clear and thorough presentation of the theoretical and mathematical aspects of the subject with examples and applications drawn from current industrial and academic research By using the computer to solve problems that include actual experimental data the author is able to cover the subject matter at a practical level The books closely integrate the theoretical chemistry being taught with industrial and laboratory practice This approach enables the student to compare theoretical projections with experimental results thereby providing a realistic grounding for future practicing chemists and engineers Each volume of Physical Chemistry includes Mathematica and Mathcad Workbooks on downloadable resources Metiu s four separate volumes Thermodynamics Statistical Mechanics Kinetics and Quantum Mechanics offer built in flexibility by allowing the subject to be covered in any order These textbooks can be used to teach physical chemistry without a computer but the experience is enriched substantially for those students who do learn how to read and write Mathematica or Mathcad programs A TI 89 scientific calculator can be used to solve most of the exercises and problems Mathematica is a registered trademark of Wolfram Research Inc Mathcad is a registered trademark of Mathsoft Engineering Education Inc

*Transport Phenomena Problem Solver*, 1984

Fundamentals of Gas Dynamics Howard W. Emmons, 2015-12-08 Volume II of the High Speed Aerodynamics and Jet Propulsion series The series which stress the more fundamental aspects of the various phenomena that make up the broad field of aeronautical science The aerodynamicist and gas dynamicist will find both the classical and the important new concepts of gas dynamics presented in an informative and stimulating manner Specialists in the study of gas dynamics have contributed Sections as follows H S Tsien The Equations of Gas Dynamics L Crocco One Dimensional Treatment of Steady Gas Dynamics A Kantrowitz One Dimensional Treatment of Nonsteady Gas Dynamics W

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TABLE OF CONTENT PART I Based on Class XI

NCERT Some Basic Concepts of Chemistry Structure of Atom Classification of Elements and Periodicity in Properties Chemical Bonding and Molecular Structure States of Matter Thermodynamics Equilibrium Redox Reactions Hydrogen s Block Elements p Block Elements Organic Chemistry Some Basic Principles and Techniques Hydrocarbons Environmental Chemistry PART II Based on Class XII NCERT The Solid State Solutions Electrochemistry Chemical Kinetics Surface Chemistry Nuclear Chemistry p Block Elements The d and f Block Elements Coordination Compounds Haloalkanes and Haloarenes Alcohols Phenols and Ethers Aldehydes Ketones and Carboxylic Acids Nitrogen Containing Compounds Biomolecules Polymers Chemistry in Everyday Life Analytical Chemistry General Principles and Processes of Isolation of Elements Questions Asked in JEE Main 2015 Solved Papers 2016 JEE Main BITSAT AP EAMCET TS EAMCET GGSIPU Solved Papers 2017 JEE Main Advanced BITSAT VIT WBJEE Solved Papers 2018 JEE Main Advanced BITSAT WBJEE Solved Papers 2019 JEE Main Advanced BITSAT WBJEE A-level Chemistry Challenging Drill Solutions (Yellowreef) Thomas Bond,Chris Hughes,2020-05-14 Full solutions to question types from top schools colleges since 2003 topical order to facilitate drilling solutions to complete and thorough encyclopedia of question types step by step solutions to trick questions tendency towards carelessness is greatly reduced most efficient method of learning hence saves time advanced tradebook complete edition eBook available visit [www.yellowreef.com](http://www.yellowreef.com) for sample chapters and more **Report of Investigations** ,1982 ,

*Physical and Chemical Equilibrium for Chemical Engineers* Noel de Nevers,2012-03-20 This book concentrates on the topic of physical and chemical equilibrium Using the simplest mathematics along with numerous numerical examples it accurately and rigorously covers physical and chemical equilibrium in depth and detail It continues to cover the topics found in the first edition however numerous updates have been made including Changes in naming and notation the first edition used the traditional names for the Gibbs Free Energy and for Partial Molal Properties this edition uses the more popular Gibbs Energy and Partial Molar Properties changes in symbols the first edition used the Lewis Randal fugacity rule and the popular symbol for the same quantity this edition only uses the popular notation and new problems have been added to the text Finally the second edition includes an appendix about the Bridgman table and its use **2024-25 NCERT Class 11th**

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## **Table of Contents Ideal Gas Law Chem 14 4 Answers**

1. Understanding the eBook Ideal Gas Law Chem 14 4 Answers
  - The Rise of Digital Reading Ideal Gas Law Chem 14 4 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ideal Gas Law Chem 14 4 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ideal Gas Law Chem 14 4 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ideal Gas Law Chem 14 4 Answers
  - Personalized Recommendations
  - Ideal Gas Law Chem 14 4 Answers User Reviews and Ratings
  - Ideal Gas Law Chem 14 4 Answers and Bestseller Lists
5. Accessing Ideal Gas Law Chem 14 4 Answers Free and Paid eBooks
  - Ideal Gas Law Chem 14 4 Answers Public Domain eBooks
  - Ideal Gas Law Chem 14 4 Answers eBook Subscription Services
  - Ideal Gas Law Chem 14 4 Answers Budget-Friendly Options
6. Navigating Ideal Gas Law Chem 14 4 Answers eBook Formats



- ePub, PDF, MOBI, and More
  - Ideal Gas Law Chem 14 4 Answers Compatibility with Devices
  - Ideal Gas Law Chem 14 4 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Ideal Gas Law Chem 14 4 Answers
    - Highlighting and Note-Taking Ideal Gas Law Chem 14 4 Answers
    - Interactive Elements Ideal Gas Law Chem 14 4 Answers
  8. Staying Engaged with Ideal Gas Law Chem 14 4 Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Ideal Gas Law Chem 14 4 Answers
  9. Balancing eBooks and Physical Books Ideal Gas Law Chem 14 4 Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Ideal Gas Law Chem 14 4 Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Ideal Gas Law Chem 14 4 Answers
    - Setting Reading Goals Ideal Gas Law Chem 14 4 Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Ideal Gas Law Chem 14 4 Answers
    - Fact-Checking eBook Content of Ideal Gas Law Chem 14 4 Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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