

## POGIL: Atomic Structure and Beginning to read the periodic table

Why?

Atoms make up everything around us—from the food on our plate to the materials that make up buildings. How do we know what to make those substances out of? How do we know what nutrients are in our food? The answer is that we know about the properties of the atoms that make them up. In order to get to that point, we must first understand the composition of atoms and how to determine that based upon the information provided to us on the periodic table.

### Model 1: Subatomic Particles

Particle	Symbol	Relative Charge	Absolute Mass	Relative Mass	Location in the atom
Electron	$e^{-}$	-1	$9.109 \times 10^{-31}$ kg	0	Electron Cloud
Proton	$p^{+}$	+1	$1.673 \times 10^{-27}$ kg	1	Nucleus
Neutron	$n^0$	0	$1.675 \times 10^{-27}$ kg	1	Nucleus

Memory Clues:

Protons: Positive

3 N's for Neutrons: Neutrons, Neutral, Nuclues

1. Draw the quantum model of the atom, labeling the three subatomic particles described above and labeling the name of their location.
2. What do you notice about the relative mass of the proton and neutron? What about the electron?

Since the masses of the subatomic particles are so small, we use another unit to discuss their mass: **atomic mass unit (amu)**. One amu is the mass of a proton. In this way, we can discuss and use the masses of protons (which, relatively speaking, is the same as the mass of a neutron) to determine the **atomic mass** of a particular atom. The atomic mass is the total number of protons + neutrons in the nucleus of the atom. The atomic mass is the average of all of the masses of the different isotopes of elements (isotopes will be discussed later).

3. Is the mass of an electron taken into account for the atomic mass? Why or why not?
4. 1 amu is equivalent to how many protons?

# Batteries Pogil Answers

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