





#### Gas Variables

How are the variables that describe a gas related?

#### What

Imagine buying a balloon bouquet at a party store. How will the beliam gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar<sup>b</sup>, which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), teregerature (T), and moles (n) of gas. These four variables can be related mathematically so that prodictions about gas behavior can be made.

#### Model 1 – Gases in a Nonflexible Container

#### Experiment A (Adding more gas)



Welliamse w. I. turning Internal pressure = 1 atm Internal pressure = 2 atm Temperature - 200 K.



Wohanne - 1 unit Temperature = 200 K.

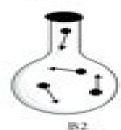


A13-Wolaume w. J. tamie External pressure = 1 ann. External pressure = 1 ann. External pressure = 1 ann. Internal pressure - 3 atm Temperature = 200 K.

#### Experiment B (Heating the gas)



Wohamse - I main Temperature = 200 K



Wollume - I unit External pressure = 1 atm. External pressure = 1 atm. External pressure = 1 atm. Internal pressure - 1 atm Internal pressure - 2 atm Internal pressure - 3 atm Temperature + 400 K



Wohamse - 1 unit Temperature = 600 K.

"Note: Volume in this model is recorded in some rather than livers because 4 molecules of gas at the conditions given would occupe a very small space (-1 × 10<sup>-15</sup> pl.). The particles shown here are much larger compared to the space boresen them obancacroad gas particles.

Cine Wariables



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**Victor M. Corman** 

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