

## GASEOUS STATE

**Measurable properties of gases :**

1. **Mass :**  
**Def.** The gases do possess mass. The mass of gas is generally used in the form of number of moles which is related as :
2. **Volume :**  
**Def.** Volume of gas is nothing but volume of the container in which it is present. Relation between different units of volume  
 $1 \text{ m}^3 = 10^3 \text{ dm}^3 = 10^3 \text{ litre} = 10^6 \text{ cm}^3 = 10^3 \text{ ml} = 10^9 \text{ mm}^3.$

3. **Temperature :**  
**Def.** Degree of hotness or coldness of a body is measured by temperature

$$\frac{C}{100} = \frac{K - 273}{100} = \frac{F - 32}{180}$$

C – Celcius scale,      K – Kelvin scale,      F – Fahrenheit scale

**Note :** In all the problems of gaseous state (i.e. in all gas law equations), temperature must be expressed in kelvin scale, i.e.,  $t^\circ\text{C} + 273 = \text{TK}$

4. **Pressure :**  
**Def.** Force acting per unit area

$$P = \frac{F}{A}$$

**Units :**

CGS :      dyne/cm<sup>2</sup>

MKS :      Newton/m<sup>2</sup>      (1N/m<sup>2</sup> = 1Pa)

Relation :      1 N/m<sup>2</sup> = 10 dyne/cm<sup>2</sup>

**Units of pressure :**

1 atm = 76 cm of Hg  
          = 760 mm of Hg  
          = 760 torr  
          = 1.01325 × 10<sup>5</sup> N/m<sup>2</sup>  
          = 101.325 kPa  
          = 1.01325 bar  
          = 14.7 lb/in<sup>2</sup> (Psi)  
          = 10.33 meters of H<sub>2</sub>O

### DENSITY OF GASES

**Absolute density**  
(mass per unit volume)

$$d = \frac{m}{V}$$

**Relative density**  
(Relative to hydrogen turned as vapour density)

$$VD = \frac{\text{molecular mass}}{2}$$

**Boyle's law and measurement of pressure :**

**Statement :**

For a fixed amount of gas at constant temperature, the volume occupied by the gas is inversely proportional to the pressure applied on the gas or pressure of the gas.

$$V \propto \frac{1}{P}$$

hence  $PV = \text{constant}$

this constant will be depend on the amount of the gas and temperature of the gas.

$$P_1 V_1 = P_2 V_2$$

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