

## Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (-)	Polyatomic Anion Name
HClO <sub>3</sub>	Chloric acid	H <sub>3</sub> O <sup>+</sup>	ClO <sub>3</sub> <sup>-1</sup>	Chlorate
H <sub>2</sub> SO <sub>3</sub>	Sulfurous acid	2 H <sub>3</sub> O <sup>+</sup>	SO <sub>3</sub> <sup>-2</sup>	Sulfite
H <sub>2</sub> SO <sub>4</sub>	Sulfuric acid	2 H <sub>3</sub> O <sup>+</sup>	SO <sub>4</sub> <sup>-2</sup>	Sulfate
H <sub>3</sub> PO <sub>3</sub>	Phosphorous acid	3 H <sub>3</sub> O <sup>+</sup>	PO <sub>3</sub> <sup>-3</sup>	Phosphite
H <sub>3</sub> PO <sub>4</sub>	Phosphoric acid	3 H <sub>3</sub> O <sup>+</sup>	PO <sub>4</sub> <sup>-3</sup>	Phosphate
HNO <sub>3</sub>	Nitric acid	H <sub>3</sub> O <sup>+</sup>	NO <sub>3</sub> <sup>-1</sup>	Nitrate
HNO <sub>2</sub>	Nitrous acid	H <sub>3</sub> O <sup>+</sup>	NO <sub>2</sub> <sup>-1</sup>	Nitrite
H <sub>2</sub> CO <sub>3</sub>	Carbonic acid	2 H <sub>3</sub> O <sup>+</sup>	CO <sub>3</sub> <sup>-2</sup>	Carbonate

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

*Ternary acids contain three elements and they also contain the element oxygen.*

b. What number do you think the prefix "ter-" refers to?

*"ter" refers to three*

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO<sub>3</sub>?

*ClO<sub>3</sub><sup>-1</sup> has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl, not HClO<sub>3</sub>.*

10. Write a rule for naming ternary acids.

*- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.*

*- "ate" is replaced with "ic" and "ite" is replaced with "ous".*

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



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**Tao Wei**



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### **Table of Contents Gas Variables Pogil Activities Answer**

1. Understanding the eBook Gas Variables Pogil Activities Answer
  - The Rise of Digital Reading Gas Variables Pogil Activities Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Variables Pogil Activities Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Variables Pogil Activities Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Variables Pogil Activities Answer
  - Personalized Recommendations

- Gas Variables Pogil Activities Answer User Reviews and Ratings
- Gas Variables Pogil Activities Answer and Bestseller Lists
- 5. Accessing Gas Variables Pogil Activities Answer Free and Paid eBooks
  - Gas Variables Pogil Activities Answer Public Domain eBooks
  - Gas Variables Pogil Activities Answer eBook Subscription Services
  - Gas Variables Pogil Activities Answer Budget-Friendly Options
- 6. Navigating Gas Variables Pogil Activities Answer eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Variables Pogil Activities Answer Compatibility with Devices
  - Gas Variables Pogil Activities Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Variables Pogil Activities Answer
  - Highlighting and Note-Taking Gas Variables Pogil Activities Answer
  - Interactive Elements Gas Variables Pogil Activities Answer
- 8. Staying Engaged with Gas Variables Pogil Activities Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Variables Pogil Activities Answer
- 9. Balancing eBooks and Physical Books Gas Variables Pogil Activities Answer
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Variables Pogil Activities Answer
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Variables Pogil Activities Answer
  - Setting Reading Goals Gas Variables Pogil Activities Answer
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Variables Pogil Activities Answer
  - Fact-Checking eBook Content of Gas Variables Pogil Activities Answer

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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